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| **4.3 Particle model of matter** |  |  |  |
| *4.3.1 Changes of state and the particle model* |  |  |  |
| **4.3.1.1 Density of materials** |  |  |  |
| a) The density of a material is defined by the equation:* $density= \frac{mass}{volume}$
* $ρ= \frac{m}{V}$

density, *ρ*, in kilograms per metre cubed, kg/m**3**mass, *m*, in kilograms, kgvolume, *V*, in metres cubed, m**3**b) The particle model can be used to explain* the different states of matter
* differences in density.

c) Recognise/draw simple diagrams to model the difference between solids, liquids and gases.d) Explain the differences in density between the different states of matter in terms of the arrangement of atoms or molecules. |  |  |  |
| **Required practical activity 5:** use appropriate apparatus to make and record the measurements needed to determine the densities of regular and irregular solid objects and liquids. Volume should be determined from the dimensions of regularly shaped objects, and by a displacement technique for irregularly shaped objects. Dimensions to be measured using appropriate apparatus such as a ruler, micrometer or Vernier callipers. |  |  |  |
| **4.3.1.2 Changes of state** |  |  |  |
| 1. Describe how, when substances change state (melt, freeze, boil, evaporate, condense or sublimate), mass is conserved.
2. Changes of state are physical changes which differ from chemical changes because the material recovers its original properties if the change is reversed.
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| *4.3.2 Internal energy and energy transfers* |  |  |  |
| **4.3.2.1 Internal energy** |  |  |  |
| 1. Energy is stored inside a system by the particles (atoms and molecules) that make up the system. This is called internal energy.
2. Internal energy is the total kinetic energy and potential energy of all the particles (atoms and molecules) that make up a system.
3. Heating changes the energy stored within the system by increasing the energy of the particles that make up the system. This either raises the temperature of the system or produces a change of state.
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| **4.3.2.2 Temperature changes in a system and specific heat capacity** |  |  |  |
| a) If the temperature of the system increases, the increase in temperature depends on the mass of the substance heated, the type of material and the energy input to the system.b) The following equation applies:* $change in thermal energy=mass x specific heat capacity$
* $ΔE=m x c Δθ$

change in thermal energy, Δ*E*, in joules, Jmass, *m*, in kilograms, kgspecific heat capacity, *c*, in joules per kilogram per degree Celsius, J/kg °Ctemperature change, Δ*θ*, in degrees Celsius, °C.c) The specific heat capacity of a substance is the amount of energy required to raise the temperature of one kilogram of the substance by one degree Celsius. |  |  |  |
| **4.3.2.3 Changes of heat and specific latent heat** |  |  |  |
| If a change of state happens:1. The energy needed for a substance to change state is called latent heat.
2. When a change of state occurs, the energy supplied changes the energy stored (internal energy) but not the temperature.
3. The specific latent heat of a substance is the amount of energy required to change the state of one kilogram of the substance with no change in temperature.
* $energy for a chnge of state=mass x specific latent heat$
* $E=m x L$

energy, *E*, in joules, Jmass, *m*, in kilograms, kgspecific latent heat, *L*, in joules per kilogram, J/kgSpecific latent heat of *fusion* is the change of state from solid to liquidSpecific latent heat of *vaporisation* is the change of state from liquid to vapour**Perform an experiment** to measure the latent heat of fusion of water.Be able to interpret heating and cooling graphs that include changes of state.Be able to distinguish between specific heat capacity and specific latent heat. |  |  |  |
| *4.3.3 Particle model and pressure* |  |  |  |
| **4.3.3.1 Particle motion in gases** |  |  |  |
| 1. The molecules of a gas are in constant random motion.
2. The temperature of the gas is related to the average kinetic energy of the molecules.
3. Changing the temperature of a gas, held at constant volume, changes the pressure exerted by the gas
4. Be able to:
* explain how the motion of the molecules in a gas is related to both its temperature and its pressure
* explain qualitatively the relation between the temperature of a gas and its pressure at constant volume.
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**PHYSICS EQUATIONS TO LEARN BY HEART**

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| **PARTICLE MODEL topic** |
|  | **Quantity** | **Unit** | **Equation** |
| *PART-ICLES* | *13* | **𝞀****m****v** | densitymassvolume | kilogram per metre3kilogrammetre3 | kg/m3kgm3 | $$ρ=\frac{m}{v}$$ |